

our data would correspond to 1,180 or 1,480 calories per mole of azeotrope, depending on whether the reaction was stipulated as involving stoichiometric quantities of the Fluothane and ether or a large excess of ether, respectively. The former of these values compares very favorably with that of Hall, Norris and Downs (1,200 calories per mole).

Finally, in response to the question raised in the aforementioned article concerning the nature of the bonding in the Fluothane-ether azeotrope, we would suggest that azeotrope formation in this system may be attributable to a hydrogen bonded structure of the form:

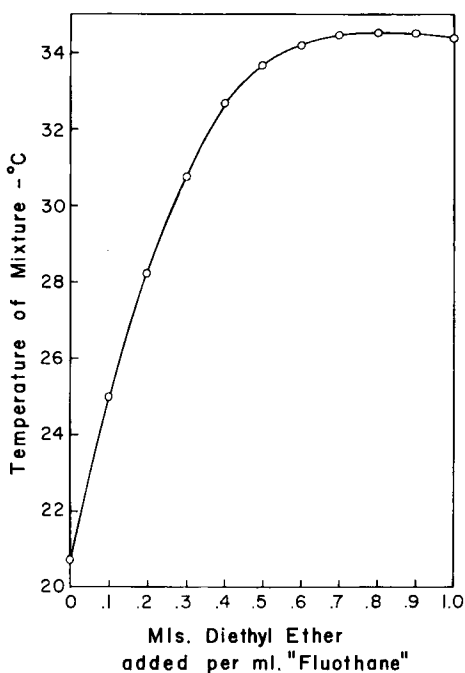
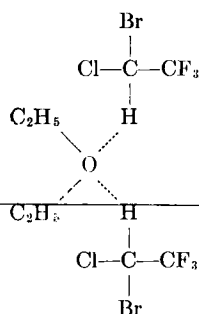


FIG. 1. Temperature rise on successive additions of diethyl ether to Fluothane.

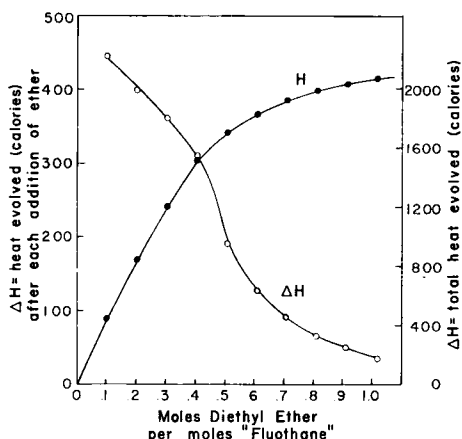


FIG. 2. Thermal data for system—Fluothane-diethyl ether: heat evolved on successive additions of diethyl ether to Fluothane.

since the oxygen atom of diethyl ether has considerable basic character (electron donor) and the hydrogen atom of the Fluothane molecule possesses enhanced acidity due to the inductive effect (high electronegativity) of the halogens, particularly the fluorine atoms.

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### Correction

To the Editor.—In my article "Concentrations of Halothane, Ether, and Cyclopropane in Inspired Atmospheres During Closed Circuit Anesthesia" (ANESTHESIOLOGY 22: 459 May-June 1961), there was an error on page 463 in the paragraph just preceding the summary of the paper. This paragraph should have read: "In the cases of halothane and ether, the concentrations in blood would probably be quite dissimilar for the two drugs due to the greater solubility of ether in blood. It is known that ether is considerably more soluble than halothane in blood and one would expect relative blood tensions of halothane to develop at a considerably faster rate than those of ether."

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