Continuous fixed bed adsorption of Cu(II) by halloysite nanotube–alginate hybrid beads: an experimental and modelling study

Yanyan Wang, Xiang Zhang, Qiuru Wang, Bing Zhang and Jindun Liu

ABSTRACT

We used natural resources of halloysite nanotubes and alginate to prepare a novel porous adsorption material of organic–inorganic hybrid beads. The adsorption behaviour of Cu(II) onto the hybrid beads was examined by a continuous fixed bed column adsorption experiment. Meanwhile, the factors affecting the adsorption capacity such as bed height, influent concentration and flow rate were investigated. The adsorption capacity \(Q_0\) reached 74.13 mg/g when the initial inlet concentration was 100 mg/L with a bed height of 12 cm and flow rate of 3 ml/min. The Thomas model and bed-depth service time fitted well with the experimental data. In the regeneration experiment, the hybrid beads retained high adsorption capacity after three adsorption–desorption cycles. Over the whole study, the new hybrid beads showed excellent adsorption and regeneration properties as well as favourable stability.

Key words | alginate, column adsorption, Cu(II), halloysite nanotube

INTRODUCTION

With the development of industry and technology, more attention has been paid to pollution caused by heavy metals which are not biodegradable and can persist and accumulate in the environment and living organisms, posing a serious threat to lives (Petrangeli Papini et al. 2004; Li et al. 2009). Copper ions especially are discharged from the industrial waste streams of the pulp, fertilizer and paint industries, water run-off, etc. Recently, various technologies have been developed to remove copper ions from effluent, such as chemical precipitation, ion exchange, adsorption, solvent extraction, membrane filtration, and electrochemical methods (Ghassabzadeh et al. 2010). However, numerous studies show that adsorption is one of the most effective alternatives for its low investment, simple operation, and potential recovery and reuse of metals.

In recent years, many biopolymer adsorbents have been widely used in the recovery of valuable metals and removal of toxic metal contaminants as they are biodegradable, environmentally friendly, efficient, and inexpensive (Hasan et al. 2006). Such materials include celluloses, alginites, carrageenans, proteins, and chitin derivatives (Seki & Suzuki 1996). In these materials, alginate mainly extracted from brown seaweed has many carboxylic groups which offer metal-binding sites and are favourable for adsorption processes. Several studies reveal that the alginate beads can be used to remove some metal ions from wastewater (Papageorgiou et al. 2008; Rocher et al. 2008). However, the use of pure alginate beads on a practical scale has been limited because of low rigidity, high mass transfer resistance and poor mechanical strength (Zhang et al. 2008). To improve the properties of alginate beads is a challenge.

Halloysite nanotubes (HNTs) with layered hollow structure and large specific surface area have been proven to be attractive adsorbents for various metal ions and organic pollutants (Luo et al. 2010; Wang et al. 2010). More importantly, HNTs can be used as reinforcing nanofibre instead of expensive carbon nanotubes in composites. Thus, the present study was devoted to fabricating halloysite nanotube–alginate (HNT-alginate) hybrid beads with high adsorption capacity and mechanical strength in aqueous solution by using halloysites as the supporting matrix. The column adsorption experiments for Cu(II) were applied to evaluate the adsorption performance of the hybrid beads. The results indicated that HNT-alginate hybrid beads could be used as...
an economical and efficient adsorbent for removal of Cu(II) from wastewater and were valuable for further application in heavy metal ion removal.

MATERIALS AND METHODS

Materials and instruments

Sodium alginate (chemical grade) was received from Paini Chemical Reagent Co. Ltd (Zhengzhou, China). HNTs were obtained by shattering and sieving clay minerals (Henan, China). Other chemicals were all analytical reagent grade and used as received. Distilled water was used in the whole process.

Using an ultraviolet and visible spectrophotometer (UV-2450, Shimadzu) with the sodium diethyldithiocarbamate spectrophotometric method, the concentration of Cu(II) in the effluent was analyzed. The morphology of hybrid beads was examined by a scanning electron microscope (SEM, JSM-7500F).

Preparation of HNT-alginate hybrid beads

The steps to prepare HNT-alginate hybrid beads were as follow. Firstly, HNTs, sodium alginate and distilled water were added to a beaker followed by stirring of the solution for 12 h at room temperature to a homogeneous mixture. Under continuous stirring, the mixture was slowly dropped into a solution of calcium chloride (2%) by syringe. Then HNT-alginate hybrid beads were formed by Ca$^{2+}$ crosslinking the sodium alginate to form hydrogel microspheres and embedding the uniformly dispersing HNTs into the microspheres. The obtained hybrid beads were left in the solution for 24 h and washed with distilled water several times and dried at 50°C for 12 h. To obtain uniform and porous beads, a series of experiments was performed and the results showed that the optimal mass ratio of HNT and sodium alginate was 1:1.

Continuous fixed bed column adsorption experiments

The continuous fixed bed adsorption experiments were conducted in a glass column with different amounts of HNT-alginate hybrid beads (Figure 1). Before operation, the bed was rinsed with distilled water and left for 6 h to ensure a closely packed arrangement of beads with no voids, channels or cracks. In down-flowed mode the Cu(II) solutions with known initial concentration and volumetric flow rate flowed through the fixed bed. The effluent solution was collected as a function of designed time at the bottom of the column. The effects of bed height, influent concentration and volumetric flow rate were studied, and all experiments were carried out at room temperature.

Regeneration of adsorbents

In the regeneration experiment 1 M HCl solution was used as an eluent. When adsorption equilibrium was achieved, the eluent was pumped into this column with the same volumetric flow rate as in the adsorption process. When the concentration of Cu(II) in the effluent solution reached zero, the desorption process finished. Then distilled water was sent through the column to remove HCl until the effluent was close to neutral. The next adsorption process was conducted under the same conditions. In this study, three adsorption–desorption cycles were performed.

RESULTS AND DISCUSSION

Figure 2(a) shows a digital photograph of the HNT-alginate hybrid beads in the wet state. It reveals that the beads possessed an average diameter of about 1.5 mm with a smooth and uniform surface, and the beads were light yellow or gray due to the presence of HNTs. To examine the fine morphology, the hybrid beads were observed by scanning electron microscopy. Figure 2(b) shows clearly that the bead was comprised of HNTs and alginate, and the HNTs overlapped loosely together in the interior of the hybrid beads. This loose structure may be more beneficial to the mass transfer of adsorbate molecules during adsorption.

The Thomas model of dynamic adsorption

A very important characteristic for an adsorption column is the breakthrough curve, whose shape determines the dynamic response and operation. The Thomas model based on Langmuir adsorption–desorption kinetics is one of the most generally used methods to describe the breakthrough curves in the study of column performance (Unuabonah et al. 2010; Barron-Zambrano et al. 2010; Song et al. 2011).
The nonlinear and linear equations of the Thomas model are shown as follows:

\[
\frac{C_t}{C_0} = \frac{1}{1 + \exp\left(K_{th}Q_0M/\nu - K_{th}C_0t\right)} \tag{1}
\]

\[
\ln\left(\frac{C_0}{C_t} - 1\right) = \frac{K_{th}Q_0M}{\nu} - K_{th}C_0t \tag{2}
\]

where \(K_{th} \text{ (ml mg}^{-1} \text{ min}^{-1})\) is the Thomas rate constant; \(C_t\) (mg/L) is the effluent concentration at any time; \(C_0\) (mg/L) is the influent concentration; \(Q_0\) (mg/g) is the maximum solid phase concentration of Cu(II) per weight of adsorbent; \(M\) (g) is the mass of the adsorbent; \(\nu\) (ml/min) is the volumetric flow rate.

According to Equation (2), the \(Q_0\) and \(K_{th}\) under different conditions, such as bed height, initial inlet concentration

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**Figure 1** | The schematic of the fixed-bed column of HNT-alginate hybrid beads for Cu(II) adsorption.
and flow rate, can be obtained from the slope and intercept of plots of ln\(\left(\frac{C_0}{C_t} - 1\right)\) versus \(t\). The values of \(R^2\), \(Q_0\) and \(K_{th}\) are listed in Table 1. The \(R^2\) of the linear plots of ln \(\left(\frac{C_0}{C_t} - 1\right)\) versus \(t\) at different conditions are all above 0.992, which suggest that the adsorption of Cu(II) onto the hybrid beads predominantly fits the Thomas model.

**Effect of bed height**

The influence of bed height was studied at different bed heights (6, 9, 12 cm) with the constant initial concentration of 50 mg/L and flow rate of 3 ml/min. All three breakthrough curves are shown in Figure 3(a) with a similar trend. As indicated in Table 1, the breakthrough time increased from 256 to 1,240 min, and the exhaustion time increased from 2,270 to 4,100 min as the bed height increased from 6 to 12 cm. And as the bed height increased, the \(Q_0\) calculated via the Thomas model increased from 46.41 to 63.07 mg/g. For the column with greater bed height has more adsorbents which can provide more adsorption sites for Cu(II). In addition, in the higher column the effluent would have enough contact time for Cu(II) to be adsorbed on the hybrid beads, while in the lower column the opposite phenomenon occurred. The result means the bed height is an important factor for the column adsorption process.

**Effect of initial inlet concentration**

The effect of initial inlet concentration was studied at different initial concentrations (50, 75 and 100 mg/L) with a bed height of 12 cm and flow rate of 3 ml/min. All three breakthrough curves are shown in Figure 3(b), and the breakthrough time, exhaustion time and the total adsorption capacity are shown in Table 1. As the table shows, with the increase of initial concentration from 50 to 100 mg/L, the breakthrough time decreased from 1,240 to 400 min, and similarly, the exhaustion time decreased from 4,100 to 2,384 min, while the \(Q_0\) increased from 63.07 to 74.13 mg/g. That may be attributed to the higher driving force provided by the greater concentration. The driving force for sorption is the concentration difference between the solute on the sorbent and in the solution (Uddin et al. 2009; Ahmad & Hameed 2010). The change of concentration gradient affects the saturation rate and breakthrough time. Therefore, the higher initial concentration provides a higher adsorption rate, and the column can reach exhaustion more quickly.

**Effect of flow rate**

To study the influence of different flow rates (3, 4.5 and 6 ml/min), the experiments were performed with a constant initial concentration of 50 mg/L and bed height of 12 cm.
The breakthrough curves are shown in Figure 3(c), and the other parameters are shown in Table 1. As expected, the breakthrough time increased from 353 to 1,240 min, and the exhaustion time increased from 2,837 to 4,100 min with the flow rate decreasing from 6 to 3 ml/min. Similarly, the $Q_0$ increased from 53.89 to 63.07 mg/g. It is considered that at higher flow rate, the adsorbates and adsorbents do not have enough contact time for interaction and Cu(II) does not have more chances to be adsorbed on the hybrid beads. While at lower flow rate, the Cu(II) can fully contact with the adsorbents. In brief, the lower flow rate the effluent has, the greater adsorption capacity the column has.

The bed-depth service time model of dynamic adsorption

The bed-depth service time (BDST) model is a simple model for predicting the relationship between bed depth and service time at a certain concentration and flow rate. The equation of the BDST model can be expressed as follows (Abu-Lail et al. 2010; Maiti et al. 2008):

$$ t = \frac{N_0 Z}{C_0 u_0} - \frac{1}{K_a C_0} \ln \left( \frac{C_0}{C_t} - 1 \right) $$

where $C_0$ (mg/L) is the influent concentration; $C_t$ (mg/L) is the Cu(II) concentration at breakthrough; $N_0$ (g/L) is the adsorption capacity of the bed; $u_0$ (cm/min) is the linear velocity; $K_a$ (L g$^{-1}$ min$^{-1}$) is the BDST rate constant; $Z$ (cm) is the bed depth; and $t$ (min) is the service time at breakthrough. According to the model, the breakthrough time can be predicted under different bed heights, which can provide reference for the column design in practical applications.

The times when $C_t/C_0$ reached 0.1, 0.2 and 0.3 were selected as column service times and Figure 4 shows the fitted curves. It can be seen that the breakthrough time
increased with the increase of bed height at the same $C_t/C_0$. As indicated in Table 2, the related coefficients almost reached 0.999, which means the results fitted well with the BDST model. As the $C_t/C_0$ increased from 0.1 to 0.3, $N_0$ increased from 24.83 to 26.92 g/L, and $K_a$ increased from 0.058 to 0.143. This indicates that the adsorption capacity per unit volume increased with the increase of $C_t/C_0$, and the amount of Cu(II) which was adsorbed on the beads also increased. So enhancing the bed height can advance the removal of Cu(II).

**Regeneration of the adsorbents**

It is critically significant for an adsorbent to possess favourable reusability in practical application. Adsorption–desorption cycle experiments were performed to evaluate the reusability of HNT-alginate hybrid beads. In this study, the column with the constant initial concentration of 50 mg/L, bed height of 9 cm, and flow rate of 3 ml/min was used for the experiments. The breakthrough curves plotted by the adsorption process of the three cycles are shown in Figure 5, and the breakthrough time, exhaust time and $Q_e$ of the three cycles are exhibited in Table 3. It can be seen that, $t_b$, $t_e$ and $Q_e$ decreased from 693 to 278 min, 3,414 to 2,561 min, and 59.59 to 38.45 mg/g respectively with the number of the cycles increasing. After three cycles, the adsorption capacity $Q_e$ still remained about 40 mg/g, which was relatively high for the adsorption of Cu$^{2+}$ ions from wastewater and indicated the good adsorption property of the prepared adsorbent. That means the hybrid beads can be regenerated and reused in practical application.

Furthermore, the experiment with three cycles had lasted more than 150 hours, and in the whole process, the adsorbents were washed by 1 M HCl more than once. Even so, the morphology and size of the hybrid beads did not change. But generally, pure alginate beads easily swell and rupture during the comparison experiments. So the HNT-alginate hybrid beads show excellent mechanical strength and stability in the aqueous solution.

In addition to good adsorption of Cu$^{2+}$ ions from wastewater, the HNT-alginate hybrid beads showed good adsorption performance for other heavy metal ions, such as Ag$^+$ ions and Cd$^{2+}$ ions.

**Table 2** The constants of the BDST model for Cu(II) adsorption on HNT-alginate hybrid beads

<table>
<thead>
<tr>
<th>$C_t/C_0$</th>
<th>$U_0$ (cm/min)</th>
<th>$C_0$ (g/L)</th>
<th>$N_0$ (g/L)</th>
<th>$K_a$ (L g$^{-1}$ min$^{-1}$)</th>
<th>$R^2$</th>
</tr>
</thead>
<tbody>
<tr>
<td>0.1</td>
<td>3</td>
<td>0.05</td>
<td>24.83</td>
<td>0.058</td>
<td>0.999</td>
</tr>
<tr>
<td>0.2</td>
<td>3</td>
<td>0.05</td>
<td>26.37</td>
<td>0.072</td>
<td>0.998</td>
</tr>
<tr>
<td>0.3</td>
<td>3</td>
<td>0.05</td>
<td>26.92</td>
<td>0.143</td>
<td>0.999</td>
</tr>
</tbody>
</table>

**Figure 4** The BDST model for adsorption of Cu(II) on HNT-alginate hybrid beads. Reaction conditions: initial concentration: 50 mg/L, flow rate: 3 ml/min.

**Figure 5** The adsorption process of the adsorption–desorption cycles. Reaction conditions: initial concentration: 50 mg/L, bed height: 9 cm, flow rate: 3 ml/min.

**ADSORPTION MECHANISM**

In this study, the main mechanism of the adsorption of Cu(II) on HNT-alginate hybrid beads can be attributed to ion exchange and complexation reactions (Chen et al. 2002; Saxena & Bajpai 2009). Alginites possess large amounts of functional groups (represented as R), such as carboxylic groups. In aqueous solution of Cu$^{2+}$ ions,
addition, there may be some free functional groups (R
with Cu
in the hybrid beads, leading to a complexation reactions
CaR
þ
is also favourable for adsorption (Luo
R2
can generate electrostatic attraction with Cu
At the same time, the surface negative charge of HNTs
exchange between Cu
ions and calcium ions inside the
hybrid beads takes place as shown in Equation (4). In
addition, there may be some free functional groups (R
in the hybrid beads, leading to a complexation reactions
with Cu
ions as shown in Equation (5):

\[ \text{CaR} + \text{Cu}^{2+} = \text{CuR} + \text{Ca}^{2+} \]  

(4)

\[ \text{R}^{2-} + \text{Cu}^{2+} = \text{CuR} \]  

(5)

Furthermore, HNTs have been proven to have a layered
hollow structure and surface negative charge (Wang et al.
2010). The layered hollow structure with large specific sur-
face is conducive to the physical adsorption of Cu
ions. At the same time, the surface negative charge of HNTs
can generate electrostatic attraction with Cu
ions, which is also favourable for adsorption (Luo et al.
2010).

### CONCLUSION

In this study, the adsorbent HNT alginate hybrid beads
showed good adsorption and regeneration properties and
were prepared by simply immobilizing alginate onto a
nanotubular structure. They also had favourable mechani-
cal and chemical stability. In the column adsorption, the
experimental data fitted well with both the Thomas and
BDST models. Meanwhile, the regeneration experimental
results showed that the adsorbent still exhibited a good
adsorption capacity after three cycles. To summarize, the
prepared HNT alginate hybrid beads have excellent
prospects for application in the wastewater treatment field.

### ACKNOWLEDGEMENT

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### REFERENCES

Abu-Lail, L., Bergendahl, J. A. & Thompson, R. W. 2010
Adsorption of methyl tertiary butyl ether on granular zeolites:
batch and column studies. Journal of Hazardous Materials

Ahmad, A. A. & Hameed, B. H. 2010 Fixed-bed adsorption of
reactive azo dye onto granular activated carbon prepared from

Barron-Zambrano, J., Szygula, A., Ruiz, M., Sastre, A. M. &
Guibal, E. 2010 Biosorption of Reactive Black 5 from aqueous
solutions by chitosan: column studies. Journal of
Environmental Management 91 (12), 2669–2675.

interactions between metal ions and Ca alginate-based ion-
exchange resin by spectroscopic analysis and modeling

Ghassabzadeh, H., Mohadespour, A., Torab-Mostaedi, M., Zaheri,
P., Maragheh, M. G. & Taheri, H. 2010 Adsorption of Ag, Cu
and Hg from aqueous solutions using expanded perlite.

Adsorption of divalent cadmium (Cd(II)) from aqueous
solutions onto chitosan-coated perlite beads. Industrial &
Engineering Chemistry Research 45, 5066–5077.

Li, J., Hu, J., Sheng, G., Zhao, G. & Huang, Q. 2009 Effect of pH,
ionic strength, foreign ions and temperature on the adsorption
of Cu(II) from aqueous solution to GMZ bentonite. Colloids
and Surfaces A: Physicochemical and Engineering Aspects
349 (1–3), 195–201.

Luo, P., Zhao, Y., Zhang, B., Liu, J. & Yang, Y. 2010 Study on the
adsorption of Neutral Red from aqueous solution onto

Maiti, A., DasGupta, S., Basu, J. K. & De, S. 2008 Batch and
column study: adsorption of arsenate using untreated laterite
as adsorbent. Industrial & Engineering Chemistry Research
47 (5), 1620–1629.

Calcium alginate beads from Laminaria digitata for the
removal of Cu
and Cd
from dilute aqueous metal

Petrangeli Papini, M., Saurini, T., Bianchi, A., Majone, M. &
Beccari, M. 2004 Modeling the competitive adsorption of Pb,
Cu, Cd, and Ni onto a natural heterogeneous sorbent
material (Italian ‘Red Soil’). Industrial & Engineering
Chemistry Research 43 (17), 5032–5041.

<table>
<thead>
<tr>
<th>Cycle</th>
<th>C0 (mg/L)</th>
<th>H (cm)</th>
<th>V (ml/min)</th>
<th>t5 (min)</th>
<th>tR (min)</th>
<th>Kd (ml g(^{-1}) min(^{-1}))</th>
<th>Qe (mg/g)</th>
<th>(R^2)</th>
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<td>1</td>
<td>50</td>
<td>9</td>
<td>3</td>
<td>693</td>
<td>3414</td>
<td>0.034</td>
<td>59.59</td>
<td>0.996</td>
</tr>
<tr>
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<td>50</td>
<td>9</td>
<td>3</td>
<td>470</td>
<td>2933</td>
<td>0.043</td>
<td>45.87</td>
<td>0.996</td>
</tr>
<tr>
<td>3</td>
<td>50</td>
<td>9</td>
<td>3</td>
<td>278</td>
<td>2561</td>
<td>0.044</td>
<td>38.45</td>
<td>0.991</td>
</tr>
</tbody>
</table>

### Table 3 | The data from column adsorption over three cycles

- **Cycle**: The cycle number of the adsorption experiment.
- **C0 (mg/L)**: The initial concentration of Cu
- **H (cm)**: The height of the column.
- **V (ml/min)**: The flow rate of the solution.
- **t5 (min)**: The time required for 5% of Cu to be adsorbed.
- **tR (min)**: The time required to reach equilibrium.
- **Kd (ml g\(^{-1}\) min\(^{-1}\))**: The distribution coefficient.
- **Qe (mg/g)**: The equilibrium adsorption capacity.
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