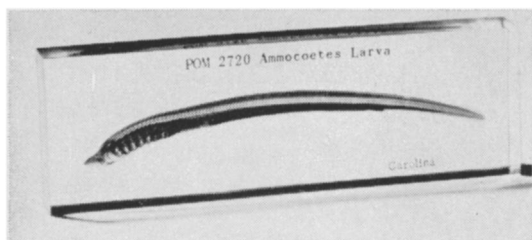


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number of interesting historical incidents that bear on the progress of the science of chemistry. The book is a useful compilation. However, the clever analogies and charming prose so characteristic of Dr. Asimov's writings are conspicuously lacking in this volume. Nevertheless, it is a useful reference volume and should be in high school libraries.

Joseph D. Novak
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EXPERIMENTS IN CHEMISTRY, John W. Sulcoski, 103 pp., \$2.50, Burgess Publishing Company, Minneapolis 15, Minnesota, 1962.

This manual contains the usual exercises found in the conventional high school laboratory and provides a good collection of concepts to be observed. It is an improvement over many on the market in that quantitative measurements and calculations are required as often as possible. Students are called on to use charts and graphs to record and interpret data and to make statements of conclusions instead of filling in blanks.

One objective of laboratory work is the improvement of techniques. The manual does this; however, very little planning is required by the student. Procedures are given in such highly detailed form that they can be followed "cook-book" fashion without much thought. No doubt the author planned it this way in order that time could be conserved for more exercises. This practice may be good for beginning experiences but could profitably be gradually eliminated as the student becomes more familiar with experimentation and thinking in terms of chemical concepts.

Virgil Heniser
Coordinator for School Science
Indiana University

ACIDS, BASES, AND THE CHEMISTRY OF THE COVALENT BOND, Calvin A. VanderWerf, 117 pp., \$7.50, Reinhold Publishing Corporation, New York, 1961.

This is one of a series of paperback books devoted to selected topics in modern chemistry. It will be useful for the high school student who is planning to continue his study of chemistry as well as for the graduate students who wish to strengthen the conceptual foundation upon which their advanced learning rests. This treatment of the covalent bond goes well beyond the quantitative aspects of chemistry and develops the qualitative principle of the relationship between structure and properties. Reactions at the covalent bond are explained in terms of free radicals and polar reactions.