Adsorption and catalytic oxidation of organic pollutants using Fe-zeolite

Analia V. Russo, Cesar Velasco Andrade, Laura E. De Angelis and Silvia E. Jacobo

ABSTRACT

In this work, a natural zeolite, modified and loaded with iron (NZ-A-Fe) as a heterogeneous catalyst, was characterized for its suitability as a permeable reactive barrier (PRB) material for treatment of aromatic hydrocarbons in groundwater. Adsorption and oxidation processes were analyzed. Batch adsorption tests for benzene, toluene and xylene (BTX) aqueous concentrated solutions were performed at neutral pH. Kinetic adsorption was described with the pseudo-second-order model. Experiments were performed using a stirred batch reactor with near 11 mM initial BTX concentration applying NZ-A-Fe as solid catalyst and H2O2 as an oxidant. BTX removal reached 80% in 600 min in these experimental conditions. Catalytic oxidation was described with a pseudo-first-order kinetic model. No significant iron leaching was detected during all the experiences. These investigations show that coupling adsorption with catalytic oxidation with this novel system is a promising procedure to simultaneously remove BTX from moderately concentrated aqueous solution at neutral pH in groundwater.

Key words | adsorption, BTX, catalyst, Fenton-like, water remediation, zeolites

INTRODUCTION

BTEX (benzene, toluene, ethylbenzene and xylene) are widespread contaminants of groundwater and soils. These volatile and water soluble chemicals may be dispersed in the environment because of accidents and storage tank leakages (Kao et al. 2006).

The physicochemical properties of the fuel oxygenate make difficult their treatment with conventional water cleaning technologies (Sutherland et al. 2004) so it is still necessary to explore different and new techniques. Adsorption process in heterogeneous phase is an economically feasible process in the removing of different classes of pollutants from water and wastewater particularly in low concentrations. The need for an alternative low-cost easily available adsorbent has encouraged the search for new adsorbents as zeolites, as an option of solid structure. They possess three basic properties: ion exchange, adsorption and molecular sieving. These properties are exploited in a wide spectrum of applications, with a global market of several million tons per year. A significant number of studies have suggested that the area of application could be expanded by functionalization of the zeolite surface. Surfactant-modified zeolites (SMZs) have been described as highly effective for adsorption in the treatment of hydrocarbon-containing wastewater, although evidencing progressive surfactant release (with negative impact both on economics and environmental protection) (Altare et al. 2007; Schick et al. 2011).

Advanced oxidation processes are a group of processes that are based on the generation of highly reactive radicals, especially hydroxyl radicals (OH*), which are extremely active and nonselective oxidants, being able to oxidize a wide range of compounds that, otherwise, can be difficult to degrade. The Fenton’s reagent (H2O2 + Fe2+) has been used to treat wastewaters containing contaminants used in industrial practices (Kang & Hoffman 1998; Neyens & Baeyens 2009; Botas et al. 2010). Oxidation by the Fenton’s reagent is traditionally represented by Equation (1). Ferric ions produced in reaction (1) can then react with H2O2

\[
\text{Fe}^{2+} + \text{H}_2\text{O}_2 \rightarrow \text{Fe}^{3+} + \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \cdot \·
to produce hydroperoxyl radicals \((HO_2^\cdot)\) and restore ferrous ions (Equation (2)).

\[
H_2O_2(ac) + Fe^{2+}(ac) \rightarrow Fe^{3+}(ac) + OH^-(ac) + H^+ + HO_2^\cdot
\]  
(1)

\[
H_2O_2(ac) + Fe^{3+}(ac) \rightarrow Fe^{2+}(ac) + HO_2^\cdot + H^+=
\]  
(2)

The hydroxyl radical \((OH^\cdot)\) is a nonselective oxidant and it can mineralize an organic compound as:

\[
C_xH_y(ac) + OH^-(ac) \rightarrow CO_2(g) + H_2O(l) + HEAT
\]  
(3)

where \(C_xH_y\) is a generic organic compound.

However, it should be pointed out that classical Fenton’s reagent, which is called as ‘homogeneous Fenton process’, leads to some significant problems related to the catalyst losses and its regeneration.

To overcome these drawbacks of homogeneous Fenton process, heterogeneous catalysts in which iron is located in the porous solid structure, have been developed. So, heterogeneous Fenton processes have been improved as a more practical and economical alternative to homogeneous Fenton process (Deng et al. 2008; Santos et al. 2010; Dükkanç et al. 2010). Nairat et al. (2015) reported the preparation of composite sorbents with iron nanoparticles and clinoptilolite for the removal of dyes, while Bayat et al. (2012) dispersed metallic clusters of iron in clinoptilolite as a catalyst for phenol degradation (initial phenol concentration: 100 mg/L and pH = 3.5) with good results. We have earlier reported (Russo et al. 2014) auspicious results for benzene removal with Fenton system at low pH, with iron nanoparticles inserted into a raw clinoptilolite (NZ-Fe). In these experiences we observed adsorption and degradation processes. Few researchers reported interesting results of organic pollutants removal (elimination) at neutral pH. Gamal et al. (2014) et al. employed ultrasonic-assisted heterogeneous Fenton reaction, using nanosized oxides of Cu and Fe, for nitrobenzene (NB) degradation where sonication played an important role in order to avoid iron precipitation.

Eun-Ju Kim et al. (2016) proposed, for wastewater treatment, manganese oxide nanorods for neutral Fenton conditions as more efficient catalysts.

In this work, the porous structure of a modified natural clinoptilolite as a supported catalyst (NZ-A-Fe) has been investigated to explore the behavior of this material in adsorption and oxidation processes. All the experiments were performed with relatively high concentrated benzene, toluene and xylene (BTX) aqueous solutions (near solubility) and at neutral pH, in order to explore the suitability of this material as a permeable reactive barrier (PBR) material for aromatic hydrocarbons treatment in groundwater.

**EXPERIMENTAL**

**Materials and methods**

Benzene (Bz) was purchased from Merck Química, toluene (Tol) was obtained from Biopack Chemical and p-xylene (Xyl) were obtained from Carboclor SA.

Two stock BTX-contaminated groundwater, with different concentrations (a and b), were prepared. They were obtained by adding the required volume (μl) of Bz, Tol and Xyl into a buffered solution to maintain pH = 7 during all the experiences.

(a) The absorption experiments were carried out using a batch operation mode with a stock solution with 1.3 mM in each component (2.3.2).

(b) For adsorption kinetic experiments with further oxidation course, a stock solution containing 5 mM Bz, 4.9 mM Tol and 1.8 mM of Xyl. was prepared.

The buffer solution was prepared adding the adequate amounts of 
Na₂HPO₄ and NaH₂PO₄ in deionized water. The neck of the flask containing the stock solution was sealed with Teflon tape and it was stirred for 24 h to complete dissolving of the organic compounds.

The concentration of freely dissolved BTX was determined by gas chromatography (GC) analysis. Iron leaching was analyzed using an atomic absorption spectrometer instrument. Temperature was fixed at 20°C for all the experiences.

**Preparation of Fe-loaded zeolite (NZ-A-Fe)**

Natural clinoptilolite (NZ) was obtained from La Rioja (Argentina) through Diaotec S.R.L. Structural characteristics were reported earlier (Russo et al. 2014, 2015a, 2015b). Cation capacity, iron content and particle size of the explored samples are presented in Table 1.

As it was earlier described, samples of this natural zeolite (NZ) were sieved and treated in NH₄Cl 3 M at 353 K during 8 h (sample NZ-A). Later, these samples were batch-loaded with iron salts (Fe(II)) and chemically reduced with sodium borohydride in order to prepare NZ-A-Fe.
Diffraction patterns and scanning electron microscopy (SEM) images of samples NZ, NZ-A and NZ-A-Fe were presented and discussed in our previous work (Russo et al. 2014, 2015a, 2015b).

Adsorption isotherms

Adsorption-desorption isotherms of NZ and the products from chemical and thermal treatments (NZ-A and NZ-A-Fe) were obtained by N2 gas adsorption using a Coulter SA3100 surface area analyzer. Samples were outgassed overnight at 140°C under vacuum for 3.0 h to attain a constant weight. The specific surface area (Brunauer–Emmett–Teller (BET)) was calculated by the BET method. \( V_{\text{tot}} \) was then evaluated by converting the volume of nitrogen adsorbed at \( p/p_0 \approx 0.98 \) to the volume of liquid adsorbate, while pore size distribution was determined from desorption branch of the isotherms by Barrett–Joyner–Halenda method (BJH). The t-plot method has been used to calculate the microspore volume.

The equilibrium sorption isotherm in aqueous solution

The equilibrium sorption isotherm in aqueous solution, were measured at 20 ± 2°C. 20 ml suspension samples with various BTX dilutions (1.3; 1.0; 0.8 and 0.5 mM) with 1.2 g NZ-A-Fe were arranged. A mixing time of 24 hours on a horizontal shaker was sufficient to reach the adsorption equilibrium. Samples were analyzed as 2.1

The BTX uptake at equilibrium (mg/g) was calculated by Equation (4).

\[
q_e = \frac{(C_0 - C_e)}{m} \cdot V
\]  

where \( C_0 \) and \( C_e \) (mg/L) are the liquid-phase concentrations of BTX at the initial and at the equilibrium phase, respectively. \( V \) is the volume of the solution (L) and \( m \) is the mass of dry adsorbent used (g)

Sorption kinetics

60-mL glass vials with screw caps with septum were used to carry out kinetic experiences BTX sample of stock solution (2.1b) were placed into the vial in contact with 3.6 g of NZ-A-Fe. The vials were placed on a horizontal shaker and aliquot of 1 ml were extracted at prefixed time. All adsorption tests were performed in duplicate and the averages are reported.

Heterogeneous Fenton reaction

Samples (from BTX stock solution 2.1b) were mixed on a horizontal shaker for 24 hours in order to reach the adsorption equilibrium. After this period, 3 ml of hydrogen peroxide (30% w/w), was added. Samples were mixed on a horizontal shaker while aliquots of 1 ml were extracted at different times. The concentration of freely dissolved BTX was determined by GC analysis. Temperature was fixed at 20°C for all the experiences.

RESULTS AND DISCUSSION

| Characteristics of the natural clinoptilolite (NZ) and of the iron loaded treated clinoptilolite (NZ-A-Fe) |
|-------------------------------------------------|-----------------|---------------|-----------------|
| Zeolite                                        | MEC (meq/g)     | EEC (meq/g)   | Particle size   | Fe content % w/w |
| NZ                                             | 1.82 ± 0.06     | 0.45 ± 0.03   | 0.3–0.5         | 1.4              |
| NZ-A-Fe                                       | –               | 0.47 ± 0.03   | 0.3–0.4         | 2.1              |

MEC, maximum experimental exchange capacity; EEC, external cation capacity.
microspores (d < 6 nm). The hysteresis loop begins at the relative pressure near 0.43 for both presented samples. The microspore volume values for raw, natural zeolite (NZ) and for the iron loaded sample (NZ-A-Fe) determined by Saito–Foley (SF) model and calculated by t-plot are presented in Table 2 showing low values (0.0021 and 0.0017 cm³/g, respectively) for both samples. The mesoporous volume, determined by BHJ model, show that the ammonium treatment followed by iron lodging decreases zeolite porosity (Table 2). The experimental porous volume values for this raw clinoptilolite (NZ) are similar to those reported by M. Sprynsky et al. for a clinoptilolite sample obtained from the Transcarpathian region in Ukraine (Sprynsky et al. 2010). Table 1 shows that iron content increases up to 50% in NZ-A-Fe compared to NZ, while the surface area decreases to a value of 9.78 m²/g mainly due to the presence of iron nanoparticles (17% lower than the surface area of NZ) while the pore volume slightly decreases (in 7% see Table 2).

Sorption isotherms

The relationships between adsorbent and adsorbate concentration at equilibrium phase can be described by sorption isotherms. The Langmuir and the Freundlich models are
the most frequently used. The Langmuir isotherm assumes monolayer coverage on a homogeneous surface with identical adsorption sites. In solution-solid systems, the isotherm adequacy can be seriously affected by the experimental conditions, in particular, the range of concentration of the solute/adsorbate.

The Freundlich isotherm is introduced as an empirical model,

\[ q_e = K_f C_e^{1/n} \]  

(5)

where \( q_e \) represents the amount adsorbed per amount of adsorbent at the equilibrium (mg/g), \( C_e \) represents the equilibrium concentration (mg/L), \( 1/n \) and \( K_f \) are parameters that depend on the adsorbate and adsorbent.

Although Langmuir and Freundlich models were explored in this work, the experimental data, results revealed that Freundlich adsorption isotherm was the best model to fit the experimental values obtained for the adsorption of BTX onto NZ-A-Fe (Figure 2(a)).

Linearizing the Equation (5) results in the following expression:

\[ \log(q_e) = \log(K_f) + \frac{1}{n} \log(C_e) \]  

(6)

where \( K_f \) and \( n \) are Freundlich constants which correspond to adsorption capacity and adsorption intensity, respectively. Figure 2(b) shows Freundlich isotherm representing the variation \( \log q_e \) with respect to \( \log C_e \) for BTX adsorption onto NZ-A-Fe.

The \( n \) value indicates the degree of nonlinearity between solution concentration and adsorption. \( n = 1 \), indicates linear adsorption; \( n < 1 \), indicates a chemical process while \( n > 1 \), indicates a physical process. Table 3 shows that the \( n \) values in Freundlich equation varies from 0.7 to 0.9, (Table 3) which shows a cooperative adsorption.

Low \( K_f \) values are related to the low adsorption capacity of this clinoptilolite, probably as a consequence of its structurally small SiO₂/Al₂O₃ ratio (near 4.3) related to hydrophobicity. Table 3 shows that \( K_f \text{Tol} > K_f \text{Bz} \) in agreement with the maximum sorption behavior of Tol on NZ-A-Fe (Table 4).

Seifi et al. (2014) reported higher values for \( K_f \) in the sequence Bz > Tol > Xyl for BTX on modified ZSM-5 while adsorption values reported by Ghiachi et al. (2004) for benzene on modified ZSM-5-88, with ratio \( 31 < \text{SiO}_2/\text{Al}_2\text{O}_3 < 88 \), were higher than our experimental results.

Sorption kinetics

Figure 3 shows the course of sorbed BTX fraction on NZ-A-Fe in the experiments for 1,800 min. BTX sorption process closely approached equilibrium within 100 min (inset Figure 3). No significant changes were observed from 120 to 1,600 min. The maximum adsorbed fraction \( (q_t) \) for Bz, Tol and Xyl were 2.7, 3.2 and 1.3 mg/g NZ-A-Fe, respectively. These \( q_t \) are in rich agreement with the \( K_f \) order obtained in Table 3 for these pollutants.

The pseudo-first-order and pseudo-second-order kinetics models were applied to analyze the kinetic and adsorption results.
The pseudo-first-order kinetic model given by Langer-green and Svenska is defined as:

\[ \ln\left(\frac{q_e}{C_0} - \frac{q_t}{C_0}\right) = \ln\left(\frac{q_e}{C_0}\right) - k_1 t \tag{7} \]

where \( q_e \) and \( q_t \) are the mg the pollutant adsorbed at equilibrium and at any time (min), respectively. This model was discarded as neither the plot of \( \ln\left(\frac{q_e}{C_0} - \frac{q_t}{C_0}\right)\) vs time followed a linear behavior (not shown) nor the calculated parameters agreed with the experimental values.

The pseudo-second-order kinetic model is expressed as:

\[ \frac{t}{q_t} = \frac{1}{k_2 q_e^2} + \frac{t}{q_e} \tag{8} \]

where \( k_2 \) (g mg\(^{-1}\) min\(^{-1}\)) is the rate constant of second-order adsorption. The linear plot (Figure 4) showed a good agreement between the experimental and the calculated \( q_e \) values (Table 4) with good correlation coefficient values for the explored concentration presented in this work.

As it can be observed in Table 4, the rate constant of the second-order adsorption, \( k_2 \) is Xyl > Tol > Bz, which indicates that the uptakes decreased with water solubility and increased with molecular weight. The reported order of uptakes on SMZ, on the other hand, was Bz > Tol > Xyl (Torabian et al. 2010). This was due to the partition mechanism, which was reversely proportional to the octanol-water partition coefficients of BTEX, where the values of B, T, E, and X were 2.13, 2.69, 3.15, and 3.15–3.20, respectively. Therefore, our explored clinoptilolite (NZ-A-Fe) could be defined as material with a moderate adsorption capacity for low water solubility compounds with high molecular weight. The \( q_e \) calculated for each pollutant is related to the amount dissolved in the aqueous phase and these values are closer to the experimental data (Figure 3) for Tol and Xyl but higher for experimental Bz data (Figure 3). Notwithstanding, the normalized standard deviation \( \Delta q_t(\%) \) shows relative low values (Table 4) compared to reported data by Tan et al. (2009). These results confirm that the pseudo-second-order kinetic model is suitable to describe the absorption kinetics.
In order to identify the adsorption mechanism, the intraparticle diffusion model proposed by Weber & Morris is tasted. (1963). It is an empirically found functional relationship where the adsorption fraction \( q_t \) varies proportionally to a function of \( t^{1/2} \)

\[
q_t = k_{pi}t^{1/2} + C_i
\]  

(9)

The initial rate of intraparticle diffusion can be obtained by linearization of the curve \( q_t = f(t^{1/2}) \) where \( k_{pi} \) is the kinetic parameter of stage i) and gives an idea about the thickness of the boundary layer. The plot \( q_t \) vs \( t^{1/2} \) is related to the mechanism involved in the adsorption process. Only if the line passes through the origin, it would be possible to accept that the intraparticle diffusion is the only rate limiting process in the explored time.

Figure 5 shows that, for near the first three hours, two different processes that can be involved. The first sharper region up to 10 min can be attributed to an external surface adsorption \( (K_{pt1}) \) while the second section where a linear behavior is observed can be related to a gradual adsorption stage where intraparticle diffusion is rate-controlled \( (K_{pt2}) \). All these processes are close dependent on the pollutants initial concentration as it was reported by Tan et al. (2009) for the trichlorophenol absorption on activated carbon.

The experimental values for the kinetic data obtained from Figure 5 are presented in Table 5. It can be noticed that \( K_{pt2} \) shows lower values than the corresponding to the first stage \( (K_{pt1}) \) as each stage is the result of different mechanism. The values of \( K_{pt2} \) are lower than those reported for 2,4,6-trichlorophenol for Tan et al. (2009) for the adsorption of an organic pollutant on activated carbon. These results can be related to the different nature and morphology of the adsorbate. Seifi et al. (2011) reported lower \( K_{pt1} \) and \( K_{pt2} \) values for the adsorption of diluted BTX aqueous solutions on surfactant-modified granulated zeolites.

The intraparticle diffusion model fits the experimental results for the explored concentrations of the pollutants (BTX) for a relatively long time interval.

### Heterogeneous Fenton reaction

Coupling adsorption with catalytic oxidation at neutral pH was motivated by the following reasons: (i) to reduce the reagents consumption, (ii) to avoid iron releasing during all the process and (iii) to improve the pollutant and catalyst contact during the adsorption process.

Heterogeneous Fenton reaction begun after the hydrogen peroxide was introduced into the reaction system. Figure 6 shows the variation of the BTX ratio \( C/C_0 \) in the aqueous solution as degradation progresses.

Figure 6(a) profiles for Bz, Tol and Xylene show similar shapes. Rapid decay of BTX was observed for the first 60 min, followed by slow removal as time progressed. Figure 6(b) shows the experimental kinetic data plotted in order to characterize the removal mechanism. The observed kinetic of degradation system was reasonably well fitted to a pseudo-first-order model:

\[
\ln C - \ln C_0 = -k_1t
\]  

(10)

where \( C \) is the concentration of the compound at any time \( t \), \( C_0 \) is the initial concentration after sorption process and \( k_1 \) is the pseudo-first-order rate constant.

In this period, BTX was removed from the solution by an oxidation reaction onto the zeolite surface where a high BTX presence is concentrated in contact with the catalyst (Fe). This phenomenon is related to the intraparticle diffusion model (Table 5) and is one of the parameters that control the removal process rate.
Kinetic results (presented in Table 6) show that all the studied pollutants have similar kinetic behavior for the degradation process for the first (60 min) and second stage (600 min) in the order Xyl > Tol > Bz.

Adsorbed BTX is removed by oxidation and new BTX migrates from the aqueous solution into zeolite. This assumption is supported by preliminary results where it was possible to observe that hydrogen peroxide concentration decreases continually when a zeolite sample, prepared as 2.4, is immersed in an hydrogen peroxide aqueous solution at pH = 7.

Solubilize iron remains below 0.15 ± 0.05 ppm during all the experiences. These results contribute to enhance the stability of the catalyst NZ-Fe prepared in this work. Such a low extend of release may be regarded as the iron resistance to the leaching process.

Taking into account the high initial concentration of BTX (near 65 mg) in contact with 3.6 g de catalyst (NZ-A-Fe), BTX removal reaches up to 80% in 600 min (near 15 mg BTX/g zeolite). These are promising results for heterogeneous Fenton processes at neutral pH for soil remediation.

Further experiences are in progress in order to take knowledge about the internal mechanism associated to the degradation processes.

CONCLUSION

The adsorption and oxidation of BTX by a natural modified clinoptilolite (SiO$_2$/Al$_2$O$_3$ near 4.3), lodged with iron was investigated. BTX adsorption for diluted samples was best fitted by Freundlich equation while kinetic adsorption for high aqueous concentrated solutions was described with the pseudo-second-order kinetic model. The intraparticle diffusion scheme fits the experimental results for a relative long time interval.

In these experiments, NZ-A-Fe demonstrates a preliminary adsorption behavior while supports the catalyst for the oxidation Fenton-like reaction.

Based on the results of this study we conclude that coupling adsorption with catalytic oxidation is a promising system to remove different organic pollutant (BTX) from aqueous solution at neutral pH without leaching of iron species. In these experimental conditions, BTX removal was near 15 mg BTX/g NZ-A-Fe. New investigations are in process in order to determine the main parameters for the design of a PRB.

ACKNOWLEDGEMENTS

The authors thank Gustavo Panizza from DIATEC S.R.L. for zeolite provision. This project was partially funded by UBACyT (2014-2017).
REFERENCES


Deng, J., Jiang, J., Lin, X., Du, C. & Xiong, Y. 2008 FeVO4 as a highly active heterogeneous Fenton-like catalyst towards the degradation of Orange II. *Applied Catalysis B: Environmental* 84, 468–473.


First received 21 September 2017; accepted in revised form 23 November 2017. Available online 7 December 2017.